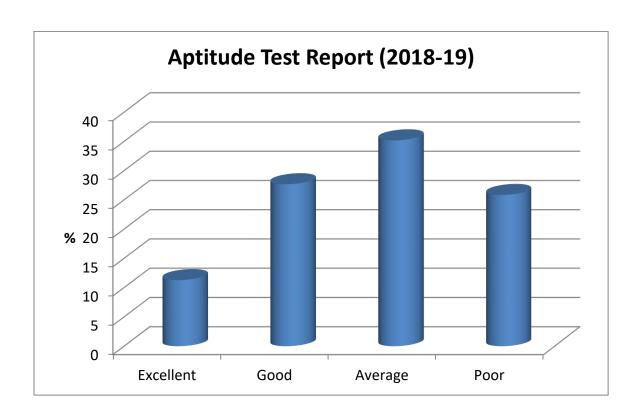
# Dr. Ambedkar College, Deekshabhoomi, Nagpur Department of Chemistry

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### **Report of Aptitude Test**

Date: 20<sup>th</sup> August 2018

The Department of Chemistry organised Aptitude test for B.Sc. SEM-I students on 20<sup>th</sup> August 2018. This aptitude test is a type of assessment that evaluates the student's talent/ability/potential to perform a basic knowledge of chemistry. Objective type questionnaire is designed based on 11<sup>th</sup> and 12<sup>th</sup> standard syllabus of chemistry. Aptitude Test results are divided in to four categories as Excellent, Good, Average and Poor. Department use aptitude test results to provide extra attentions towards average and poor category students. The aptitude test was conducted and evaluated by Prof. N. G. Telkapalliwar, Dr. V. M. Shivankar and Dr. Mrs. D. M. Borikar, Teacher In-charge of B.Sc, SEM-I.



# **Photographs during the Aptitude Test**







# **Question Paper of Aptitude Test**

	DR. AMBEDKAR COLLEGE, DEEKSHABHOOMI, NAGPUR.
	DEPARTMENT OF CHEMISTRY
	APTITUDE TEST
	(Academic Session-2018-19)
N	ame: Date:
*	********************************
	1. The mass less particles are
	a) Alpha rays b) beta rays c) gamma rays d) protons
	2. The energy production into the sun and the stars is due-to
	a) Nuclear fission b) nuclear fusion c) thermo chemical reaction d) none of these
	3. The dipole moment is maximum for
	a) HCI b) HBr c) HI (A) HF
	4. The type of hybrid orbital used by boron in BF <sub>3</sub> are
	a) Sp Ue) sp <sup>2</sup> c) sp <sup>3</sup> d d) sp <sup>3</sup>
	5. Which of the following products is not derived from cellulose?  a) Reyon bi Insulin c) Gun cotton d) paper
	a) Neyon (a) msuin (c) dun cotton (d) paper
	6. A real gas most closely approaches the behaviour of a perfect gas under the conditions of
	(a) high P and low T b) high P and high T c) low P and low T d) low P and high T
	7. The compressibility factor of ideal gas is
	a) zero (b) one c) infinity d) less than one
	8. The third law of thermodynamics stated that in the limit T→0
	a) G=0 b) H=0 c) U=0 \d) S=0
	9. For an adiabatic process
	a) T=Constant \b) q=0 c) q=constant d) W=0
	10. Isotonic solutions have
	a) the same freezint point b) the same boiling point c) the same surface tension
	the same osmotic pressure
	11. the carbon atoms in a benzene ring are
	a) sp hybridized b) sp <sup>3</sup> hybridized $(x)$ sp <sup>2</sup> hybridized d) none of these
	12. the pH of an aqueous solution of 10 <sup>8</sup> M solution of HCl is
	a) 8 (b) Slightly less than 7 c) slightly greater than 7 d) 6

12 Flores and the second secon	
13. Electromagnetic radiation with minimum wavelength is	
a) ultraviolet (b) X-rays c) Infrared d) radiowaves	
14 When the source and the	
14. When the source emitting lines is placed in a strong magnetic field the dpecra	al lines are split
a) Compton effect (b) Zeeman effect (c) Rydberg effect (d) Photoelectr	ic effect
15. The degenerate orbitals have	
Layer Equal energy h) same orientation is	
Equal energy b) same orientation in space c) same size d) all o	f the above
16. After filling the 4p orbitals, an electron will enter in	
a) 4d b) 4f Vc) 5s d) 3d	
17 occupy the same place in the periodic table	
a) Isobars b) isotopes c) isotones d) none of these	
dy notice of these	
18. The atom of contains no neutrons	
hydrogen b) deuterium c) tritium d) none of these	
19. Which form of radioactivity is most penetrating?	
a) alpha particles b) beta particles c) neutrons d) gamma rays	
20. In N <sub>2</sub> molecule these are	
a) one sigma and one pi bond b) tow sigma and one pi bond	
one sigma and two pi bond  d) three sigma bond	* *
21. in sn <sup>3</sup> sn <sup>2</sup> and sn bub side at 1	
21. in sp <sup>3</sup> , sp <sup>2</sup> and sp hybridized carbon atom the p character is maximum in	
21. in sp <sup>3</sup> , sp <sup>2</sup> and sp hybridized carbon atom the p character is maximum in sp <sup>3</sup> sp <sup>3</sup> b) sp <sup>2</sup> c) sp d) all have equal p character	
c) sp d) all have equal p character	
22. Homolytic fission leading to the formation of	
c) sp d) all have equal p character	
22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions	
22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions  23. The pH of neutral solution is	
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22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions  23. The pH of neutral solution is  a) 5 b) 7 c) 8 d) 10	
22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions  23. The pH of neutral solution is a) 5 b) 7 c) 8 d) 10  24. Maximum no. of electrons that a d-subshell contains are	
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22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions  23. The pH of neutral solution is a) 5 b) 7 c) 8 d) 10  24. Maximum no. of electrons that a d-subshell contains are a) 6 b) 2 c) 14 contains are a) 6 b) 2 c) 14 contains are	
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22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions  23. The pH of neutral solution is a) 5 b) 7 c) 8 d) 10  24. Maximum no. of electrons that a d-subshell contains are a) 6 b) 2 c) 14 cd 10  25. In living system most of the Nitrogen is found in the form of a) free nitrogen and prostation of a letter and a l	
22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions  23. The pH of neutral solution is a) 5 b) 7 c) 8 d) 10  24. Maximum no. of electrons that a d-subshell contains are a) 6 b) 2 c) 14 cd 10  25. In living system most of the Nitrogen is found in the form of a) free nitrogen at Maximum of the Nitrogen is found in the form of	
22. Homolytic fission leading to the formation of free radical b) carbocation c) carbanion d) ions  23. The pH of neutral solution is a) 5 b) 7 c) 8 d) 10  24. Maximum no. of electrons that a d-subshell contains are a) 6 b) 2 c) 14 cd 10  25. In living system most of the Nitrogen is found in the form of a) free nitrogen and prostation of a letter and a l	

### **Student List with Marks**

Dr. Ambedkar College, Deekshabhoomi, Nagpur Department of Chemistry

#### **Aptitude Test**

Session: 2018-19

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Class: B.Sc. SEM- I Date of Examination

20 August 2018 Marks: 2

Name of Teacher Incharge: 1) N. G. Telkapalliwar

2) Dr. Mrs. D. M. Borikar .....

3) Dr. V. M. Shivankar

Indina?

Sr. No.	Name of Student	Signature	Marks Obtained	Remarks
1	Ambaush . D. Dwivedi	Amberia	12	
2	Aditya Zade	-Kode	11	
3	Rituile Wikey	Wilking	15	
4.	Himanshy . V. Sukhdeve	100	13	
5	Yamesh S. Lilhare	amoh lilhar	18	
6	Harshal Bambole	Paul	17	
7.	Ravi Madko	Knedte.	16	,
8.	Vedant Sautam	Vel	16	
9.	Tarun Palil		18	
10.		glanding	12	
1)	Tromson Janbhulkan	-AT	17	
12.	Sanghasish N. UKey	Sanukey,	15	
13.	Savyatah B. Meshyam	Baus	09	
14.	Moyors. Ghazat	Jaminores	07	
15 .	Jahir N. Sheikh	Twheets	13	
	Mikhil y Borkur	BOXKOZ	09	
17.	Shreyash A. More	Breyash	12	
18.	Abhinay Akash A. Mah to	AbhinaV	13	
	Sarthak R. Parate	Daharate	16	
20.	Ashutosh D. Ambache	Ambades	15	
	Tanmay Deshmukh	Ches Communication of the Comm	12	

22.	VIVEK . P. MESHRAM	Gink	14	
23.	Samit R. Dhoke	SR Dhore	09	
24.	Shivam R. Newad	hivan	09	
	Abhinay 5. Chauchaus	Aus	- 11	
	Ashika · A · Tambe	Ashiker	07	
	Supriya . S. Patil	Fatil	08	
28.	Anuskka Jumde	mushka	16	
29.		GAncyoni.	15	
30 -	Prasidelhi Thakre	Pshake	10	
31.	Anjali Patil	(Aug Palit	15	
32.	yayri Ravit	Quit!	13	
33.	Chetana P. Gautam	Printain	18	
840	Gashrita Kanile ke	Parite	17	
35	Janua Bose	or or	16	
36	Hitika Rathi	Beteka	18	
37	Ekta Wadishasme	Elta.	16	
348	Anchal Banche	appandue	18	
39	Bhagyashree Dhuldhar.	Muldhaz	17	
40.	Sonali . G. Diyemale	Digewale	17	
41.		Sabhase	12	
42.	Ankita Shewade	whenate_	16	
43.	Vaibhan B. Patil	Partil.	10	,
44.	Sulmia R. Gailmead	Quikwad.	09	
45.		10 Bliggoth	12	
46	Acual N. evekelo	Alwerde	17	
47	Popia singh	Engl.	17	
48	Pooja singh Banjana S. Laddhad	Bladdhad	17	
49	Manadi M. Sabne	Manasi Calone,	19	
50.	Bhupali Kalita	Bhupali	18	
51-	Pragati M. Mure,	Pragati	22	
52-		Brighnau.	18	
53-		A.R.Dandeisa)	13	
54	Ratiti R. Patil	Cratiti.	10	
55	Sayali A. Gaibhiye	shajbniye	15	
56	Sayali A. Gajbhiye drya Malature	Americate .	16	
57	Shectal · Good	Seetal	to	

58	Deepak A. Prajapati Vivek R. Meshram Psaiche N. Raikway	Deepak	20	
59	Vivek R. Meshram	vivel.	18	
60	Poceche N. Raikuray	Bucksi	10	
61	Acland M. Valenton	Agarmen	12	
62	Chagydship o pandykar fran y a th Telang Parikshif W. Gtorpad Shalab Tiwasi	At reallburg	15	
63.	Pragnya H. Telana	Plag	16	
64	Paritshit W. Ghowad	70	16	
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## Dr. Ambedkar College, Deekshabhoomi, Nagpur

#### **Department of Chemistry**

### **Aptitude Test**

Session: 2018-19

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Class: B.Sc. SEM- I Da

Date of Examination

Name of Teacher Incharge: 1) N. G. Telkapalliwar

2) Dr. Mrs. D. M. Borikar

3) Dr. V. M. Shivankar

Sr.	Name of Student	Signature	Marks	Remarks
No.			Obtained	
1)	Priyal Manoj Dhunde	Dhunde	17	
2)	Prajwal Ganvir	Ring	15	
3)	vishal so Jachar	Bladbal	15	
4>	Nikhil. D. Bhole	Tetale.	06	
5)	Dhammpal- 3. Wasnik	Devisnik	05	
6)	ALTAN. 1. GUPTO	Engn.	14	
7)	rolayur. S. Tinjoude	(other w)	17	,
	Dwidnesh & Kharikar	avilly.	16	
9)	Archit. Cy. Fambhule	Archie	19	
101	Priyash 5. Bambal	Plantael	13	
11)	Sudhanshy Gire	and.	17	
/	Dudhusk Bobade	Kulled	16	
	Dhawal Shende	District Control of the Control of t	15	
14)	Sujyor Meshroam	Sugo	15	
15)	Suran sheware	Juan.	15	
	Kanchan Shimpi	Janimpi.	17	
16)	SHRYAMI MPSHYa	- Ann	14	
187	Рична С. Дрире	Rhypl	07	
19)		stuge.	09	
20)	Minha S. Pathan	Merathan.	15	
21>		Muriole	. 13	

22) Natasha : Chowdhuy	Natarhe	17	
21) Mansi Narnaware	A CONTRACTOR	20	
22) Sakshi Ghodke	Shodke.	()	
23) Sakshi kalamkay	Skalamkay	17	
24) Apeksha Hedaoo	invedo.	17	
110	C JEN W	17	
25) ShrishH Maitsa 26) Pragati Sharma	Varine	01	
27 Ketki Narnaware	Domanare.	11	
28 PRANJALI PATIL	Ritis	12	
29) Khushboo lautue	Khushloo.	15	
2 0	Bonalou.	11	
	Ford.	12	
		16	
32) Alshwayya Sudhan	Myor	18	
33) TANAYA MUDE 34) Shreya vikas Pande	forande	16	
	Bighé.	18	
35) Swalali P. Dighe	dulia.	12	
36) Aboli Tayawade 37) Gerayat Mohad	bhui	15	
	Brunka	15	
38) Bhunika Kanathe	Bhagat.	10	
32) Brandi Bhagat	Rides.	16	
40) sidra Kunain	stekaul	17	4
41) Sakshi Tekom	Buya	16	
427 Spriya Vishwakayma	Inaway.	09	
<13) Shunshite Nataghere	Skennaki	06	
44) Shoushali Kannakle	Shul	12	
45) Sueta Maghmare.	Unu		
Cal Daudilla Tumani	7/100	17	
46) Dayshita Jumane	Thuman	12	
47) Illita Khillare	2 Same	20	
47) Ishita Khilari	Duyah	20	
47) Likita Khilar. 48) Bhagyashue K. Bawankule 49) Ankita S. Thakore	Alature °5	20 20	
47) Ishita Khilar. 48) Bhagyashue K. Bawankule 49) Ankita S. Ihakee 50) Samiksha D. Wankar	Mature of	20 20 19 19	
47) Ishita Khilari 48) Bhagyashue K. Bawankule 49) Ankita S. Thakore 50) Somiksho D. Hankar 51) Shraddha Ratnakar	Alature °5	20 20 19 19 12	
47) Ishita Khilar.  48) Bhagyashue K. Bawaykule  49) Anhita S. Jhakre  50) Samiksha D. Hankar  51) Shraddha Ratnakar  52) Mahi Ipadeo	Mature of	20 20 19 19 12	
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41) Ishita Khilari 48) Bhagyashue K. Bawankule 49) Ankita S. Ihakee 50) Samiksha D. Wankar 51) Shraddha Ratnakar 52) Mahi upadeb 53) Shrawhu Kaut 54) Passavie S. Bhadake	Abstrace of the state of the st	20 20 19 19 12 13 14	
44) Ishita Khilar. 48) Bhagyashue K. Bawankule 49) Ankita S. Ihakee 50) Samiksha D. Wankar. 51) Shraddha Ratnakar 52) Mahi upadeb 53) Shrawhu Raut	Abtue "s Ference "s	20 20 19 19 12 13	

56.	Prantali V. Sherki	Dhodii	04	
57.	Pranjali V. Sherki Neha S. Boukau	Aboelkale	07	
58.	Shravani. H. Khobragade	<b>1</b> .	10	
	Kartiki Jonnalwar	( Jonnales es	10	
60.		Athabragood	12	
61.	Ayush Byogshiye	At habragood	08	
62.	Vaishnavi Hankay	Quarkas	20	
	Trusha Karwade	Hamado.	15	
63.	Mamta Paurcle	Mamboo	16	
64		Blue.	12	
	Saba Khan	The inur	07	
66		Brill	16	
	Sanchita Shelke	2-1-11/2-	09	
67)	Lydia Mary Saxrana	C J-0		
68)	Anjali Pardhi	myatr.	13	
70)	Anjali rarani	amanda	17	
71)	Uma Singh Aachal B. Hadler	200	05	
72)	Jeet Chaudhari		08	
1	Jeet Challand		13	
73	Mayur Barkers Praywal Grufter		16	
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